

Nomenclature Test Review

Name _____

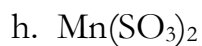
Nomenclature Questions:

- List the molecular prefixes in order 1 – 10.
 - State when you use these prefixes in naming compounds.
 - When do you NOT use a prefix?
- Write the formula for the following substances from their names:
 - diphosphorus tetraoxide
 - carbon tetrachloride
 - dinitrogen trioxide
 - carbon monoxide
 - trinitrogen decachloride
 - pentasulfur heptabromide
- Give the name of the following compounds:
 - S_4N_2
 - SiS_2
 - NO
 - P_2O_5
 - SF_6
 - N_3O_8
- Write the oxidation number with the correct sign (ion charge) for the following substances:
 - Lithium
 - Iron
 - Silver
 - Copper (I)
 - Aluminum
 - Iron (III)
- Write the formulas of the following compounds from their names:
 - Sulfur trioxide
 - Dinitrogen tetroxide
 - Trinitrogen monoxide
 - Tetraphosphorus hexachloride
 - Disulfur pentafluoride
 - Heptasulfur decaiodide

6. Name the following compounds from their formulas:



7. Name the following compounds from their formulas. (Hint think about bond type first!)



8. **State the error(s) in the following names of compounds, and how to fix it.**

1. Tin hypochlorite

4. Zinc (II) hydroxide

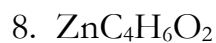
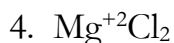
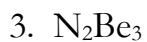
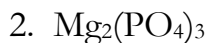
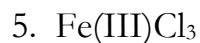
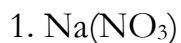
2. Ammonium oxygen

5. Iron nitrate

3. Sulfate aluminum

6. Lithide carbonate

9. **Correct the error(s) in the following formulas, and how to fix it.**



Decide if the following compounds are molecular or ionic. Then, name the compound.

- | | | | |
|-----------------------------------|----------------------|----------------------------------|---------------------|
| 1. CO ₂ | 3. CaBr ₂ | 5. NO ₂ | 7. H ₂ O |
| 2. Fe ₂ O ₃ | 4. AlCl ₃ | 6. P ₃ O ₅ | 8. PbO ₂ |

Give the formulas for the following compounds (there are ionic and molecular):

- | | |
|-------------------------|---------------------------|
| 1. carbon tetrachloride | 6. diphosphorus pentoxide |
| 2. magnesium chloride | 7. iron (III) chlorate |
| 3. iron (III) oxide | 8. magnesium phosphide |
| 4. sodium sulfate | 9. magnesium phosphate |
| 5. sulfur dioxide | 10. trisulfur dioxide |

Give the names for the following compounds (there are ionic and molecular):

- | | |
|----------------------|---------------------------------------|
| 1. Na ₂ O | 6. Ca(ClO) ₂ |
| 2. NO ₃ | 7. Mg(HCO ₃) ₂ |
| 3. CO | 8. SO ₂ |
| 4. FeBr ₂ | 9. P ₂ O ₅ |
| 5. FeBr ₃ | 10. Cu ₂ S |

Name the following compounds.

- | | | |
|-----------------------|-----------------------------------|------------------------------------|
| 1. KCl | 5. Na ₂ S | 9. Au ₃ PO ₄ |
| 2. CuMnO ₄ | 6. Fe ₂ S ₃ | 10. Ag ₃ N |
| 3. NiCl ₂ | 7. MgO | 11. Ca(OH) ₂ |
| 4. Cu ₂ O | 8. FeBr ₃ | 12. SnO ₂ |

If given a name, provide the formula. If given the formula, provide the name.

- | | |
|-------------------------|-----------------------------------|
| 1. sulfur trioxide | 11. AlF ₃ |
| 2. carbon monoxide | 12. SCl ₃ |
| 3. dinitrogen tetroxide | 13. MgSO ₃ |
| 4. calcium chloride | 14. CuMnO ₄ |
| 5. iron (II) oxide | 15. NiCl ₂ |
| 6. aluminum oxalate | 16. ZnHPO ₄ |
| 7. magnesium nitride | 17. Fe(CN) ₃ |
| 8. aluminum chloride | 18. P ₄ O ₈ |
| 9. iron (III) sulfite | 19. Na ₂ O |
| 10. hydrogen peroxide | 20. Cu ₃ N |