

**Ionic Nomenclature****Directions:** Name the following compounds.

1.  $\text{CaCO}_3$  \_\_\_\_\_
2.  $\text{KCl}$  \_\_\_\_\_
3.  $\text{FeSO}_4$  \_\_\_\_\_
4.  $\text{LiBr}$  \_\_\_\_\_
5.  $\text{Al(OH)}_3$  \_\_\_\_\_
6.  $\text{Fe}_2\text{O}_3$  \_\_\_\_\_
7.  $\text{Zn}_3(\text{PO}_4)_2$  \_\_\_\_\_
8.  $\text{SrF}_2$  \_\_\_\_\_
9.  $\text{CuC}_2\text{H}_3\text{O}_2$  \_\_\_\_\_
10.  $\text{PbO}$  \_\_\_\_\_

**Directions:** Write the formulas for the following names

1. Magnesium Nitrite \_\_\_\_\_
2. Nickel (III) Bromide \_\_\_\_\_
3. Sodium Phosphide \_\_\_\_\_
4. Silver Sulfite \_\_\_\_\_
5. Manganese(II)Hydroxide \_\_\_\_\_
6. Aluminum Acetate \_\_\_\_\_
7. Tin (IV) Sulfide \_\_\_\_\_
8. Cesium Nitride \_\_\_\_\_
9. Ammonium Chloride \_\_\_\_\_
10. Iron (III) Oxide \_\_\_\_\_

**Covalent Nomenclature****Directions:** Name the following compounds

1.  $\text{SO}_2$  \_\_\_\_\_
2.  $\text{NO}_2$  \_\_\_\_\_
3.  $\text{PCl}_5$  \_\_\_\_\_
4.  $\text{NH}_3$  \_\_\_\_\_
5.  $\text{OF}$  \_\_\_\_\_
6.  $\text{C}_2\text{Cl}_4$  \_\_\_\_\_
7.  $\text{P}_2\text{O}_7$  \_\_\_\_\_
8.  $\text{N}_2\text{Cl}_3$  \_\_\_\_\_
9.  $\text{SF}_6$  \_\_\_\_\_
10.  $\text{CO}$  \_\_\_\_\_

**Directions:** Write the compounds for the following names.

1. Carbon disulfide \_\_\_\_\_
2. Xenon tetrafluoride \_\_\_\_\_
3. Phosphorus trichloride \_\_\_\_\_
4. Octachloride monoxide \_\_\_\_\_
5. Trinitrogen heptafluoride \_\_\_\_\_
6. Dihydrogen monoxide \_\_\_\_\_
7. Pentasulfur hexacarbide \_\_\_\_\_
8. Silicon dioxide \_\_\_\_\_

**Naming Compounds (All Mixed Up!)**

**Directions:** Name the following compounds; remember they can be either ionic or covalent so be sure to use the correct rules.

1. CuCN \_\_\_\_\_
2. NiCO<sub>3</sub> \_\_\_\_\_
3. BCl<sub>4</sub> \_\_\_\_\_
4. CBr<sub>7</sub> \_\_\_\_\_
5. Zn(OH)<sub>2</sub> \_\_\_\_\_
6. Ag<sub>3</sub>N \_\_\_\_\_
7. MnS \_\_\_\_\_
8. Li<sub>3</sub>PO<sub>4</sub> \_\_\_\_\_
9. N<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

**Directions:** Write the compounds for the following names. Remember they can be either ionic or covalent so be sure to use the correct rules.

1. Ammonium phosphate \_\_\_\_\_
2. Calcium chloride \_\_\_\_\_
3. Diphosphorous pentoxide \_\_\_\_\_
4. Lead (IV) chromate \_\_\_\_\_
5. Zinc Nitrate \_\_\_\_\_
6. Strontium bromide \_\_\_\_\_
7. Tetrachlorine dibromide \_\_\_\_\_
8. Copper (II) sulfate \_\_\_\_\_
9. Iron (II) oxide \_\_\_\_\_
10. Hexasulfur trifluoride \_\_\_\_\_

**Naming Acids****Directions:** Name the following acids.

1.  $\text{H}_3\text{PO}_4$  \_\_\_\_\_
2.  $\text{HCl}$  \_\_\_\_\_
3.  $\text{H}_2\text{SO}_4$  \_\_\_\_\_
4.  $\text{HNO}_2$  \_\_\_\_\_
5.  $\text{HI}$  \_\_\_\_\_
6.  $\text{HC}_2\text{H}_3\text{O}_2$  \_\_\_\_\_
7.  $\text{H}_2\text{CO}_3$  \_\_\_\_\_
8.  $\text{HBr}$  \_\_\_\_\_

**Directions:** Write the compound for the following acids.

9. Chlorous acid \_\_\_\_\_
10. Sulfuric acid \_\_\_\_\_
11. Acetic acid \_\_\_\_\_
12. Hydrocyanic acid \_\_\_\_\_
13. Hydroselenic acid \_\_\_\_\_
14. Carbonic acid \_\_\_\_\_
15. Perchlorous acid \_\_\_\_\_
16. Bromic acid \_\_\_\_\_

**Nomenclature Errors**

**Directions:** Identify the errors in the names or compounds below. Be sure to correct the error on the line provided.

1. aluminum (III) iodide \_\_\_\_\_

2.  $\text{Al}_3\text{O}_2$  \_\_\_\_\_

3.  $(\text{NH}_4)_3\text{Cl}_2$  \_\_\_\_\_

4. lead (II) oxygen \_\_\_\_\_

5. monocalcium dioxide \_\_\_\_\_

6.  $\text{K}(\text{ClO}_2)$  \_\_\_\_\_

7.  $(\text{OH})_3\text{Al}$  \_\_\_\_\_

8.  $\text{Cr}_4(\text{CN})_3$  \_\_\_\_\_

9. copper (I) chloric \_\_\_\_\_

10.  $\text{NiNO}_{32}$  \_\_\_\_\_

11.  $\text{Mg}_2\text{F}$  \_\_\_\_\_

12. ClK \_\_\_\_\_

13.  $\text{Ca}_2\text{O}_2$  \_\_\_\_\_

14. sodium (I) fluoride \_\_\_\_\_

15. magnesium dihydroxide \_\_\_\_\_