Name: _____ Date: _____ Block: _____

Using the Chemistry EOC Reference Tables

Activity series: Determine if the following single-replacement reactions occur. If the reaction does not occur, write DNR. If a reaction occurs, write a balanced equation for the reaction.

1. $Ca + Mg(NO_3)_2 \rightarrow$
2. $Zn\ (s) + NaOH\ (aq) \rightarrow$
3. $K\ (s) + ZnCl_2\ (aq) \rightarrow$
4. $Cl_2\ (g) + HF\ (aq) \rightarrow$
5. $Cl_2\ (g) + HBr\ (aq) \rightarrow$

Solubility Rules: Complete the reactions below. Identify the precipitate formed as a result of the double replacement reactions below. If no precipitate is formed, write *DNR*.

➢ If the compound formed is insoluble in water, it will form a precipitate (solid).

1. $AgNO_3\ (aq) + NaCl\ (aq) \rightarrow$
2. $CaCl_2\ (aq) + Na_2CO_3\ (aq) \rightarrow$
3. $Cu(NO_3)_2\ (aq) + Na_2S\ (aq) \rightarrow$

   Write the net ionic equation:

4. $Al_2(SO_4)_3 + 3Mg(OH)_2 \rightarrow$

   Write the net ionic equation:

Use your Reference Tables to predict the products of these reactions:

1. $LiHCO_3 \rightarrow$
2. $MgO + H_2O \rightarrow$
3. $C_2H_6 + O_2 \rightarrow$

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