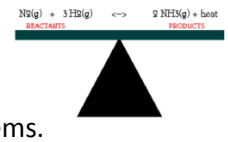
What's Your Stress? – Le Chatelier Lab

Le Chatelier's Principle
"Restoring Balance"



If you add more to the left side (*Reactants*), the lever will tip to the left. In order to restore balance, the system responds by making more of what is on the right (*Products*) and vice versa.

Purpose: You will use the following website, http://www.harpercollege.edu/tm-ps/chm/100/dgodambe/thedisk/equil/equil.htm in order to analyze three equilibrium systems.

Procedure:

Part 1:

1. Start at the **Introduction** and then proceed through the **Background** portion of the website.
2. Complete the **pre-lab questions** on your paper. There are four questions total.
3. Proceed to the **Experiment** section of the website.
4. Click on General Comments about the Experiment and read the directions for each specific part of your data table.
5. Go back to the original website and work through the example.
6. Once you have finished with the example, choose **THREE** equilibrium systems to analyze for your data table below. You should analyze **THREE** stresses for **EACH** equilibrium system.
7. Once you have finished your experiment, complete the **post-lab questions**. There are only four of them.

Data:

Equilibrium System	Stresses	Shifts	Reactant Concentrations	Product Concentrations	Keq

Answers to Pre-lab Questions:

1. _____
2. _____
3. _____
4. _____

Answers to Post-lab Questions

1. _____
2. _____
3. _____
4. _____